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on the periodic table to your right 3. On the arrows on the diagram, indicate the trend for ionization energy as you move down a column and across a row. (Options: easily loses an electron OR holds on tight to electrons) 4. Of all the elements on the periodic table, identify & circle the element with the highest & lowest ionization energy. 5.

ANSWER KEY Trends on the Periodic Table

Periodic Trends Worksheet Rank the following elements by increasing atomic radius: carbon, aluminum, oxygen, potassium. Rank the following elements by increasing electronegativity: sulfur, oxygen, neon, aluminum. What is the difference between electron affinity and ionization energy? Cc Why does fluorine have a higher ionization energy than iodine?

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Compared to the atomic radius of a sodium atom, the atomic radius of a magnesium atom is smaller. The smaller radius is primarily a result of the magnesium atom having (A) a larger nuclear charge (B) a smaller nuclear charge (C) more principal energy levels (D) fewer principal energy levels ____26.

Name Worksheet: Periodic Trends Period

Answers for Comparing Atomic Sizes Here are answers for the questions above. a. Li, C, F All are in the same period and thus have the same number of energy levels. Therefore, the important factor is the nuclear charge. Li is the largest because it has the smallest nuclear charge and pulls the ... Microsoft Word - Periodic Trends Worksheet Author:

Periodic Trends Worksheet

S-C-5-3_Periodic Trends Worksheet and KEY ANSWER KEY Periodic Trends Worksheet 1. Using the data below, make a bar graph of atomic radius vs. atomic number for Group 2A and for Period 3 of the periodic table. Group 2A Element Atomic Number Be 4 Mg 12 Ca 20 Sr 38 Ba 56 Atomic Radius 1.11 1.60 1.97 2.15 2.17.

S-C-5-3_Periodic Trends Worksheet and KEY | Periodic Table ...

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